

T08D01 – (8.2) Properties of Acids and Bases Notes

Name

1. 8.2.1 Outline the characteristic properties of acids and bases in aqueous solution. (2)

Properties of Acids	Properties of Bases

- a. What is an electrolyte?
- b. What happens to a base in water?

	What happens?	Relevant Equations:	In general:
Acid with Metal			
Acid with Metal Carbonates			
Acid with Bases			
Acid with Metal Oxides			
Acid with Metal Hydroxides			

- c. In general, provide a diagram (web) for the generic reactions acids are able to undergo:

- d. Use the example of hydrogen chloride (s) dissolved in water and the organic material methyl benzene to describe why water is important to the function of acids:

Test	Solution of HCl in water	Solution of HCl in C ₆ H ₅ CH ₃
Universal Indicator Paper		
Addition of CaCO ₃		
Conductivity		
Enthalpy of Solution		

- e. What is the difference between “hard” and “soft” water?

- f. What is the difference between a “normal salt” and an “acid salt?”